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Chapter 11 -- Organometallic

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and

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chlorine gas react to form solid  
iron(III)chloride.  $\text{CuS (s) + O}_2\text{(g) Cu}$   
 $\text{(s) + SO}_2\text{(g) D}$

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Guide      chemical equation A  
representation of a chemical reaction

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with reactants on the left, products on the right, and an arrow separating the two skeleton

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Lesson Objectives • Describe chemical reactions using word equations. • Know the correct symbols to use in order to write skeleton equations for chemical reactions. • Use coefficients to balance chemical equations so that the law of conservation of mass is

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Reactions. The reactants are written  
on the left and the products on the  
right. The arrow that separates them  
is called yield. Reactants Products.



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Symbols in Equations. Symbol

Meaning " yields D reversible reaction  
(s) solid (l) liquid (g) gas (aq) aqueous  
" catalyst " heat. Pt .

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reaction that has the general form  $AB$   
 $A + B$ . Cleavage reactions are the

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reverse of condensation reactions.),  
and oxidation–reduction reactions A  
chemical reaction that exhibits a  
change in the oxidation states of one  
or more elements in the

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Question 1. Give an example of a complex chemical reaction. Answer:

‘ Photosynthesis ’ is a complex chemical reaction. Question 2. Give an example of simple chemical reactions.

Answer: Rusting of iron and

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combustion of fuels are simply  
chemical reactions. Question 3.

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chemical equation. skeleton equation.  
catalyst. coefficient. A short, easy way

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to show a chemical reaction, using symbols. a chemical equation that does not indicate the relative amount.... a substance that initiates or accelerates a chemical reaction....

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To describe a chemical reaction. To calculate the quantities of compounds produced or consumed in a chemical reaction. As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate



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balanced equation. A \_\_\_\_\_  
is an equation where each side of the  
equation has the same number of  
atoms of each element, and mass is

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conserved. word equation. Iron + oxygen  $\rightarrow$  iron (III) oxide is an example of a(n) \_\_\_\_\_ .  
chemical equation.  $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$  is an example of a(n) \_\_\_\_\_ .

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Reading Answer Key. coefficient. a whole number that appears before a formula in an equation. spectator ion. a particle not directly involved in a chemical reaction. combustion reaction. a reaction in which oxygen reacts with another substance, often producing light or heat. reactant.

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## Chapter 11 Chemical Reactions Answers

Stoichiometry Problems. When we carry out a reaction in either an industrial setting or a laboratory, it is easier to work with masses of substances than with the numbers of molecules or moles. We will first

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Present the method used in most other books for converting from the mass of any reactant or product to the mass of any other reactant or product using a balanced chemical equation outlined ...

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