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class-12, unit-7 Chemistry Review Sheet Unit 7~~
are representative of each type of problem worked in this
unit. For more practice on a given type of problem, refer
back to Note Taking Guides, worksheets, and quizzes. 1.
Calculate the molar mass for ethane, C₂H₆. C₂H₆ 2(12.0
g) + 6(1.0 g) = 30.0 g 2. Convert 45 g Se to moles of Se. ?
mol Se = 45 g Se ! 1 mol Se 79.0 g Se = 0.57 mol Se 3.

~~Review Sheet: Unit 7 Name Vocabulary: Fill in the blanks ...~~
are representative of each type of problem worked in this
unit. For more practice on a given type of problem, refer
back to Note Taking Guides, worksheets, and quizzes. 1.
Calculate the molar mass for ethane, C₂H₆. 2. Convert 45 g
Se to moles of Se. 3. Convert 4.77 x 10²⁴ molecules of SO₂
to grams.

~~Review Sheet: Unit 7 Name Vocabulary: Fill in the blanks ...~~
UNIT 7 REVIEW SHEET – EQUILIBRIUM • EQUILIBRIUM is
the point at which the rate of the forward reaction is equal to
the rate of the reverse reaction in a reversible process. o
Reaction is dynamic – still occurring – even though to the
eye, it looks like nothing is changing at equilibrium. o
Observable/measurable quantities (concentration, pressure,
color) are constant at equilibrium. o The amounts of
products and reactants are typically not equal at equilibrium!

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-u2013...~~

CHEMISTRY CP Name: _____ Per: _____ TEST: _____ Unit 7
Review Sheet: Thermochemistry KEY Potential Energy
Diagrams: Choose the appropriate letter. 1. What is the H
of the forward reaction? f 2. What is the activation energy of
the forward reaction? b 3. What is the H of the reverse
reaction? f 4.

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semester of college chemistry, all the terms, units of

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Review Sheet: Unit 5 Name_____ CHEMISTRY: A Study of Matter © 2004, GPB 5.24a KEY I. For each of the following statements, write " I " for ionic, " C " for covalent

~~CHEMISTRY: A Study of Matter~~

Review Sheet for Chemistry Unit 5 December 11, 2011 31.
Balance the following equations by putting the proper coefficients wherever they are needed. a) $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$ b) $\text{Al} + \text{H}_2\text{O} \rightarrow \text{Al}(\text{OH})_3 + \text{H}_2$ c) $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ d) $\text{C}_5\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ e) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{H}_2 + \text{NaOH}$ f) $\text{NaClO}_3 \rightarrow \text{NaCl} + \text{O}_2$ g ...

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