

Electron Probability Lab Answers

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Instead, the quantum-mechanical model for the atom gives the probabilities of finding an electron in a particular region around the nucleus. In this experiment, we model the probable locations of electrons around the nucleus. We will use dried beans or peas to represent the electrons to help us visualize the regions of low and high density.

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Electron Probability Lab Answers Electron Structure—1s Orbital—Super Value Laboratory Kit Section 9: Electron Configuration and Quantum Numbers [ELECTRON PROBABILITY LAB] Graham Mueller Hour 8 Purpose The purpose of this lab is to investigate the probability distribution of marks about a central point.

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According to modern atomic theory, we cannot be sure of the exact location of electrons in an atom. We predict that electrons will be relatively close to the nucleus (because the electrons are negatively charged and the nucleus is positively charged).

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N=3, L=1, M=1, M=1/2. This electron is found within the 2p orbital, which has the dumbbell shape, primarily based on the z axis. This electron would share this orbital with another electron, but would have a negative spin.

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(d) If we count the number of dots in each spherical shell, we obtain the total probability of finding the electron at a given value of *r*. Because the surface area of each shell increases more rapidly with increasing *r* than the electron probability density decreases, a plot of electron probability versus *r* (the radial probability) shows a peak. This peak corresponds to the most probable radius for the electron, 52.9 pm, which is exactly the radius predicted by Bohr's model of the hydrogen ...

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