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thermal reservoir at constant temperature T_h , converts an amount W of this heat into work, and expels the remainder Q_c into another thermal reservoir at a constant temperature T_c . The efficiency e of the heat engine is the fraction of Q_h that is converted into work, so $e = W/Q_h = (Q_h - Q_c)/Q_h = 1 - Q_c/Q_h$ (1) Since the absorbed heat Q

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This agrees with Schroeder's Figure 1.14. To put in some numbers, we can use the data from the earlier problem. For water, the values given by Schroeder are (at 25 C = 298 K): = $2.57 \times 10^4 \text{ K}^{-1}$ (20) $T = 4.52 \times 10^{10} \text{ Pa}^{-1}$ (21) For one mole of water, the volume is $18.068 \times 10^{-6} \text{ m}^3$, so $C_P - C_V = (298) \times 18.068 \times 10^{-6} \times 2.57 \times 10^4 - 4.52 \times 10^{10} \times 18.068 \times 10^{-6} \times 10^{-6}$ (22) = 0.787 J K^{-1} (23) C

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